



## Allenylporphyrins: a new motif on the porphyrin periphery

Oliver B. Locos, Katja Dahms, Mathias O. Senge\*

School of Chemistry, SFI Tetrapyrrole Laboratory, Trinity College Dublin, Dublin 2, Ireland

### ARTICLE INFO

#### Article history:

Received 30 November 2008

Revised 3 March 2009

Accepted 12 March 2009

Available online 18 March 2009

### ABSTRACT

Palladium-catalysed Suzuki–Miyaura cross-coupling proved to be the simplest and most efficient method for attaching an unsubstituted allene, a novel substituent and potential functional group, to the porphyrin periphery. Comparative studies on various allenyl synthetic pathways show that this method, which makes use of a bromoporphyrin and allenylboronate, affords the corresponding allenylporphyrin in the presence of an appropriate base and THF.

© 2009 Elsevier Ltd. All rights reserved.

The synthesis and functionalisation of porphyrins are of interest to organic chemists. Tetrapyrroles play pivotal, diverse roles in Nature<sup>1</sup> and there is a continuing desire to exploit their properties for applications in medicine,<sup>2</sup> catalysis<sup>3</sup> and nano materials.<sup>4</sup> As a result, functional group interconversion and synthetic transformations spanning the field of porphyrin chemistry are constantly being explored and improved.<sup>5</sup> One niche in the area of synthetic transformations that appear to have received little to no attention with porphyrins, however, is the use of 1,2-propadiene, or allene.

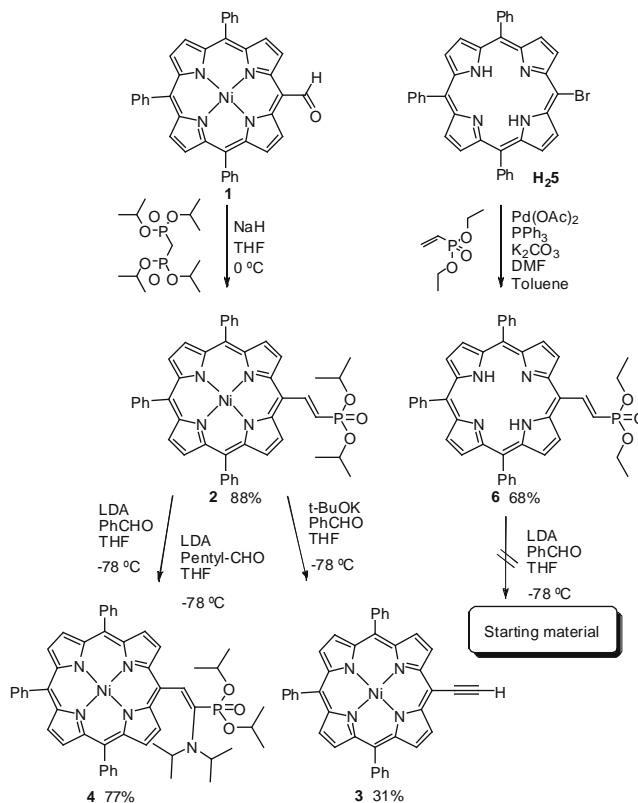
Allenes represent a versatile functional group that can be utilised as a building block in a variety of synthetic transformations.<sup>6</sup> With the emergence of efficient protocols for their preparation, allenes have allowed chemists to access a variety of structurally interesting products that possess biological, chiral and optical activity.<sup>7</sup>

While cumulenyl porphyrin dimers linked by two carbons have been explored for their impressive optical properties,<sup>8</sup> these dimers are quinoidal in nature, which results in exceptionally perturbed electronic absorption spectra. To the best of our knowledge, no studies have been performed on the utilisation of cumulative double bonds on porphyrins for their chiral and biological aspects, hence this has become the aim of our investigation.

Of the methods we envisaged being successful for allene synthesis on the macrocycle periphery, the Horner–Wadsworth–Emmons (HWE) reaction<sup>9</sup> appealed the most due to its versatility in allene generation<sup>9</sup> and to its need for a formylporphyrin precursor. Formylporphyrins are commonly used precursors the utility of which has recently been expanded successfully from solely metalloporphyrins to include their metal-free analogues.<sup>10</sup>

Applying the one-pot double olefination procedure designed by Tomioka and co-workers<sup>11</sup> to formylporphyrin **1** (Scheme 1), the HWE reaction using tetraisopropyl methylenebisphosphonate and NaH worked well and the extremely polar intermediate **2**

could be isolated by simple chromatography in very good yield. Interestingly though, both in the one-pot procedure, and using the isolated intermediate **2**, the second olefination reaction could not be realised under a variety of conditions. This olefination



Scheme 1.

\* Corresponding author. Tel.: +353 1 896 8537; fax: +353 1 896 8536.  
E-mail address: [sengem@tcd.ie](mailto:sengem@tcd.ie) (Mathias O. Senge).

reaction occurs through nucleophilic substitution of an in situ generated carbanion intermediate by an aldehyde.<sup>11</sup> However, in our case, when using either benzaldehyde or hexanal in conjunction with LDA, the result was the unexpected substitution of diisopropylamine from the base onto the alkenephosphonate to generate porphyrin **4**. Reaction of diisopropylamine with **2** occurred seemingly regardless of the temperature used ( $-80\text{ }^{\circ}\text{C}$  to  $60\text{ }^{\circ}\text{C}$ ). Use of the non-nucleophilic base, <sup>t</sup>BuOK, only resulted in partial dehydrophosphonation to yield the alkynylporphyrin **3** in modest amounts when the reaction mixture was heated to  $60\text{ }^{\circ}\text{C}$  after addition of the base. Using the metal-free analogue **6**, which was synthesised by Heck coupling, only starting material was retrieved from the reaction mixture.

As the double HWE reaction did not afford the desired result, we tried a more unconventional synthetic route to allenes by using the allenylboronic acid pinacol ester reagent and Suzuki coupling (Scheme 2). Previously, this particular reagent has been used mainly to take advantage of the allene in three- and four-component reactions and in Ru<sup>II</sup> catalysis to generate alkenylboronates.<sup>12,13</sup>

All Suzuki coupling reactions were attempted using a ten-fold excess of the allenylboronate, and the reaction mixture was degassed by three freeze-pump-thaw cycles before being heated to the desired temperature. Using a variety of conditions to couple the allenylboronate to the bromoporphyrin, it was discovered that very few conditions resulted in the consumption of the bromo starting material. However, when successful, depending on the

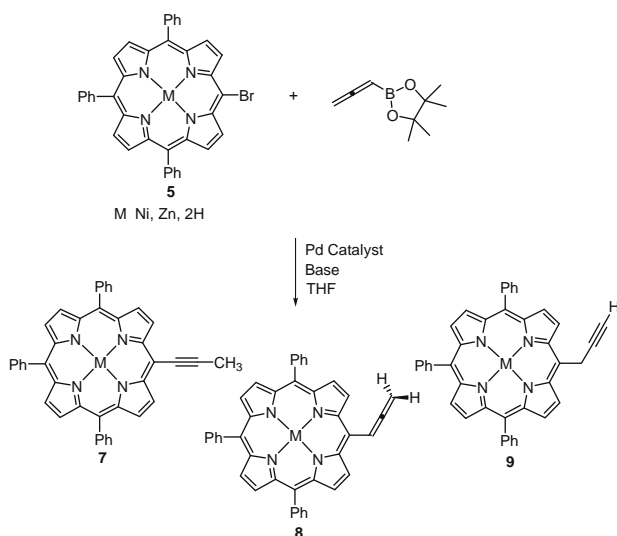
base and metalloporphyrin used, either the 1-propynyl-, **7**, allenyl-, **8**, or 2-propynylporphyrin, **9** could be isolated after column chromatography. The results of the investigations to find the optimum conditions for attaching the allene to the porphyrin are shown in Table 1.

Except for when  $\text{K}_3\text{PO}_4$  was used as base, the Suzuki coupling was relatively unsuccessful when carried out at reflux in THF. Even using  $\text{K}_3\text{PO}_4$ , coupling only occurred with the Ni<sup>II</sup> bromoporphyrin, and the use of this with a strong base resulted in rearrangement of the allene to the more stable 1-propynylporphyrin Ni<sup>II</sup> **7**, along with the formation of unidentifiable side products. Our first successful attempt at isolating the allenylporphyrin Ni<sup>II</sup> **8** came from heating in THF to  $80\text{ }^{\circ}\text{C}$  in a sealed Schlenk tube and using a 20-fold excess of  $\text{K}_3\text{PO}_4$  (9% yield). Porphyrin Ni<sup>II</sup> **7** was also isolated from this reaction in 14% yield along with unidentifiable side products. Interestingly, using the same conditions, but in 1,4-dioxane at  $80\text{ }^{\circ}\text{C}$ ,  $100\text{ }^{\circ}\text{C}$  and  $120\text{ }^{\circ}\text{C}$ , no allenyl- or propynyl-porphyrin was isolated.

On varying the Pd catalyst, the best coupling results were achieved with Pd<sup>II</sup> salts with the bidentate ligands 1,2-(diphenylphosphino)ethane and 1,3-(diphenylphosphino)propane. In particular, the allenyl-porphyrin Ni<sup>II</sup> **8** was isolated in up to 50% yield when  $\text{K}_2\text{CO}_3$  was used as the base in 10-fold excess.<sup>14</sup> Interestingly, while  $\text{PdCl}_2(\text{dppe})$  and  $\text{PdCl}_2(\text{dppp})$  were found to be equally effective in coupling, when a twofold excess of  $\text{Cs}_2\text{CO}_3$  was used, the  $\text{PdCl}_2(\text{dppp})$  caused  $\sim 25\%$  rearrangement from the allene to the 1-propynylporphyrin (entry 7). The use of  $\text{PdCl}_2(\text{dppe})$  gave no such rearrangement, so it was considered the catalyst of choice for the remaining optimisation experiments. It should be noted that the Pd<sup>II</sup> catalysts were synthesised according to the procedure of Miyaura and Suzuki,<sup>15</sup> and these represent new applications for the catalysts  $\text{PdCl}_2(\text{dppe})$  and  $\text{PdCl}_2(\text{dppp})$ .

The amount of base, strength of base and type of metalloporphyrin used in the coupling reactions were also found to have a profound effect on the outcome of the reaction. Using a twofold excess of  $\text{Cs}_2\text{CO}_3$  resulted in Ni<sup>II</sup> **8** solely being isolated from the reaction, while a 10-fold excess resulted in Ni<sup>II</sup> **7** being the only Suzuki product in 61% yield.<sup>14b</sup>  $\text{Cs}_2\text{CO}_3$  was found to be an inappropriate base when either a Zn<sup>II</sup> or free-base porphyrin was used, the result being unidentifiable products. Using the less labile alkali base,  $\text{K}_2\text{CO}_3$  in 10-fold excess led to Ni<sup>II</sup> **8** again being isolated as the only product from Ni<sup>II</sup> **5**, while quite surprisingly, the kinetically rearranged Zn<sup>II</sup> **9** was the only product isolated from the reaction of Zn<sup>II</sup> **5**.<sup>14c</sup>  $\text{K}_2\text{CO}_3$  was ineffective for porphyrin H<sub>2</sub>**5**, while  $\text{Na}_2\text{CO}_3$  did not aid in Suzuki coupling for any of the metalloporphyrins.

Inspired by the base-catalysed kinetic rearrangement of the allene in the Zn<sup>II</sup> porphyrin to form Zn<sup>II</sup> **9**, the same rearrangement was attempted for Zn<sup>II</sup> **7** (Scheme 3). Zn<sup>II</sup> **7** was prepared by a method adapted from Newman-Evans et al. for phenylprop-2-yne from phenylacetylene,<sup>16</sup> which subjects the alkyne to a strong base with



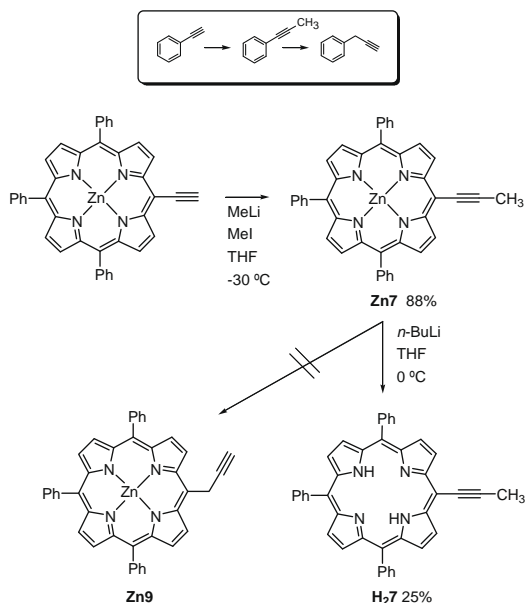
Scheme 2.

**Table 1**  
Yields of the Suzuki–Miyaura coupling reaction between **5** and allenylboronate in THF

Entry	M	Catalyst	Cat. Conc. (% mol)	Time (h)	Base (excess)	Temp $^{\circ}\text{C}$	Yield <b>7</b> (%) <sup>a</sup>	Yield <b>8</b> (%) <sup>a</sup>	Yield <b>9</b> (%) <sup>a</sup>
1	Ni	$\text{Pd}(\text{PPh}_3)_4$	15	15	$\text{Cs}_2\text{CO}_3$ (twofold)	80	—	2	—
2 <sup>b</sup>	Ni	$\text{Pd}(\text{PPh}_3)_4$	15	15	$\text{K}_3\text{PO}_4$ (twofold)	67	25	—	—
3 <sup>b</sup>	Ni	$\text{Pd}(\text{PPh}_3)_4$	15	15	$\text{K}_3\text{PO}_4$ (20-fold)	80	14	9	—
4	Ni	$\text{PdCl}_2(\text{PPh}_3)_2/\text{AsPh}_3$	25	15	$\text{Cs}_2\text{CO}_3$ (twofold)	80	—	8	—
5	Ni	$\text{PdCl}_2(\text{PPh}_3)_2/\text{AsPh}_3$	15	5	$\text{Cs}_2\text{CO}_3$ (twofold)	80	—	10	—
6	Ni	$\text{Pd}_2(\text{dba})_3/\text{AsPh}_3$	15	15	$\text{Cs}_2\text{CO}_3$ (twofold)	80	—	10	—
7	Ni	$\text{PdCl}_2(\text{dppp})$	15	15	$\text{Cs}_2\text{CO}_3$ (twofold)	80	9	37	—
8	Ni	$\text{PdCl}_2(\text{dppe})$	15	15	$\text{Cs}_2\text{CO}_3$ (twofold)	80	—	41	—
9	Ni	$\text{PdCl}_2(\text{dppe})$	15	15	$\text{Cs}_2\text{CO}_3$ (10-fold)	80	61	—	—
10	Ni	$\text{PdCl}_2(\text{dppe})$	15	15	$\text{K}_2\text{CO}_3$ (10-fold)	80	—	50	—
11	Zn	$\text{PdCl}_2(\text{dppe})$	15	15	$\text{K}_2\text{CO}_3$ (10-fold)	80	—	—	46

<sup>a</sup> Isolated yield.

<sup>b</sup> The strength of the base resulted in another porphyrin product that was not identifiable by <sup>1</sup>H NMR and MS analysis.



subsequent trapping of the anion intermediate with MeI. However, only starting material and the demetallated **H27** were retrieved from the reaction mixture after column chromatography when subjecting 2-propynylporphyrin **Zn7** to 2 equiv of *n*-BuLi in THF at 0 °C. This preliminary investigation is ongoing in our laboratory with Ni<sup>II</sup> derivatives.

In conclusion, for the first time, unsubstituted allenes and propargyl substituents have been attached to the porphyrin periphery by Suzuki coupling and trapping of an alkyne anion with MeI. The phenomenon of allene rearrangement on the porphyrin periphery is still being investigated. However, as allenes are particularly reactive functional groups, especially towards metal catalysis,<sup>7,17,18</sup> we have also begun exploring their reactivity on the macrocycle.

## Acknowledgements

This work was generously funded by a Science Foundation Ireland Research Professorship Award (SFI 04/RP1/B482) and we used facilities provided by the Centre of Chemical Synthesis and Chemical Biology funded by the HEA.

## References and notes

- (a) *The Porphyrin Handbook*; Kadish, K. M., Smith, K. M., Guillard, R., Eds.; Academic Press: San Diego, California, 2000; Vol. 1, (b) Milgrom, L. R. *The Colors of Life: An Introduction to the Chemistry of Porphyrins and Related Compounds*; Oxford University Press: London, 1997.
- (a) Maillard, P.; Loock, B.; Grierson, D. S.; Laville, I.; Blais, J.; Doz, F.; Desjardins, L.; Carrez, D.; Guerquin-Kern, J. L.; Croisy, A. *Photodiagn. Photodyn. Ther.* **2007**, *4*, 261; (b) Molinari, A.; Bombelli, C.; Mannino, S.; Stringaro, A.; Toccaceli, L.; Calcabrini, A.; Colone, M.; Mangiola, A.; Maira, G.; Luciani, P.; Mancini, G.; Arancia, G. *Int. J. Cancer* **2007**, *121*, 1149; (c) Kiesslich, T.; Berlanda, J.; Plaetzer, K.; Krammer, B.; Berr, F. *Photochem. Photobiol. Sci.* **2007**, *6*, 619; (d) Moghissi, K.; Dixon, K. *Photodiagn. Photodyn. Ther.* **2005**, *2*, 135; (e) Kessel, D.; Dougherty, T. J. *Rev. Contemp. Pharmacol.* **1999**, *10*, 19; (f) Wiehe, A.; Shaker, Y. M.; Brandt, J. C.; Mebs, S.; Senge, M. O. *Tetrahedron* **2005**, *61*, 5535.
- (a) Aviv, I.; Gross, Z. *Chem. Eur. J.* **2008**, *14*, 3995; (b) Zhang, H.-J.; Liu, Y.; Lu, Y.; He, X.-S.; Wang, X.; Ding, X. *J. Mol. Catal. A: Chem.* **2008**, *287*, 80; (c) Korotchenko, V. N.; Severin, K.; Gagne, M. R. *Org. Biomol. Chem.* **2008**, *6*, 1961; (d) Chatterjee, D. *Coord. Chem. Rev.* **2008**, *252*, 176.
- (a) Ren, D.-M.; Guo, Z.; Du, F.; Liu, Z.-F.; Zhou, Z.-C.; Shi, X.-Y.; Chen, Y.-S.; Zheng, J.-Y. *Int. J. Mol. Sci.* **2008**, *9*, 45; (b) Nobukuni, H.; Shimazaki, Y.; Tani, F.; Naruta, Y. *Angew. Chem., Int. Ed.* **2007**, *46*, 8975; (c) Kojima, T.; Nakanishi, T.; Harada, R.; Ohkubo, K.; Yamauchi, S.; Fukuzumi, S. *Chem. Eur. J.* **2007**, *13*, 8714;

- (d) Langa, F.; Gomez-Escalonilla, M. J.; de la Cruz, P. J. *Porphyrins Phthalocyanines* **2007**, *11*, 348.
- For recent examples of porphyrin functionalization, see: (a) Senge, M. O.; Richter, J. J. *Porphyrins Phthalocyanines* **2004**, *8*, 934; (b) Sugiura, K.-i.; Kato, A.; Iwasaki, K.; Miyasaka, H.; Yamashita, M.; Hino, S.; Arnold, D. P. *Chem. Commun.* **2007**, 2046; (c) Esdaile, L. J.; Jensen, P.; McMurtrie, J. C.; Arnold, D. P. *Angew. Chem., Int. Ed.* **2007**, *46*, 2090; (d) Atefi, F.; Locos, O. B.; Senge, M. O.; Arnold, D. P. *Porphyrins Phthalocyanines* **2006**, *10*, 176; (e) Baba, H.; Chen, J.; Shinokubo, H.; Osuka, A. *Chem. Eur. J.* **2008**, *14*, 4256; (f) Hori, T.; Peng, X.; Aratani, N.; Takagi, A.; Matsumoto, T.; Kawai, T.; Yoon, Z. S.; Yoon, M.-C.; Yang, J.; Kim, D.; Osuka, A. *Chem. Eur. J.* **2008**, *14*, 582; (g) Senge, M. O.; Fazekas, M.; Notaras, E. G. A.; Blau, W. J.; Zawadzka, M.; Locos, O. B.; Ni Mhuircheartaigh, E. M. *Adv. Mater.* **2007**, *19*, 2737; (h) Senge, M. O. *Acc. Chem. Res.* **2005**, *38*, 733; (i) Horn, S.; Sergeeva, N. S.; Senge, M. O. *J. Org. Chem.* **2007**, *72*, 5414; (j) Horn, S.; Senge, M. O. *Eur. J. Org. Chem.* **2008**, 4881; (k) Hao, E. H.; Fronczek, F. R.; Vicente, M. G. H. *Chem. Commun.* **2006**, 4900; (l) Fan, D. Z.; Taniguchi, M.; Lindsey, J. S. *J. Org. Chem.* **2007**, *72*, 5350; (m) Mori, G.; Shinokubo, H.; Osuka, A. *Tetrahedron Lett.* **2008**, *49*, 2170; (n) Maes, W.; Ngo, T. H.; Vanderhaeghen, J.; Dehaen, W. *Org. Lett.* **2007**, *9*, 3165.
- (a) Castro, A. M. M. *Chem. Rev.* **2004**, *104*, 2939; (b) Sydnes, L. K. *Chem. Rev.* **2003**, *103*, 1133; (c) Mikami, K.; Shimizu, M. *Chem. Rev.* **1992**, *92*, 1021; (d) Nakai, T.; Mikami, K. *Chem. Rev.* **1986**, *86*, 885; (e) Smadja, W. *Chem. Rev.* **1983**, *83*, 263; (f) Taylor, D. R. *Chem. Rev.* **1967**, *67*, 317; (g) Pasto, D. J. *Tetrahedron* **1984**, *40*, 2805; (h) Krause, N.; Hoffmann-Roeder, A. *Tetrahedron* **2004**, *60*, 11671; (i) Hoffmann-Roeder, A.; Krause, N. *Angew. Chem., Int. Ed.* **2002**, *41*, 2933.
- (a) Ma, S.; Hou, H.; Zhao, S.; Wang, G. *Synthesis* **2002**, 1643; (b) Hoffmann-Roeder, A.; Krause, N. *Angew. Chem., Int. Ed.* **2004**, *43*, 1196.
- (a) Blake, I. M.; Rees, L. H.; Claridge, T. D. W.; Anderson, H. L. *Angew. Chem., Int. Ed.* **2000**, *39*, 1818; (b) Piet, J. J.; Taylor, P. N.; Anderson, H. L.; Osuka, A.; Warman, J. M. *J. Am. Chem. Soc.* **2000**, *122*, 1749; (c) Beljonne, D.; O'Keefe, G. E.; Hamer, P. J.; Friend, R. H.; Anderson, H. L.; Bredas, J. L. *J. Chem. Phys.* **1997**, *106*, 9439.
- (a) Nagaoka, Y.; Tomioka, K. *J. Org. Chem.* **1998**, *63*, 6428; (b) Nagaoka, Y.; Inoue, H.; Tomioka, K. *Phosphorus, Sulfur Silicon Relat. Elem.* **2002**, *177*, 1843; (c) Wadsworth, W. S., Jr.; Emmons, W. D. *J. Am. Chem. Soc.* **1961**, *83*, 1733; (d) Fillion, H.; Refouvet, B.; Pera, M. H.; Dufaud, V.; Luche, J. L. *Synth. Commun.* **1989**, *19*, 3343; (e) Fillion, H.; Hseine, A.; Pera, M. H. *Synth. Commun.* **1987**, *17*, 929.
- (a) Runge, S.; Senge, M. O. *Tetrahedron* **1999**, *55*, 10375; (b) Dahms, K.; Senge, M. O.; Bakar, M. B. *Eur. J. Org. Chem.* **2007**, 3833; (c) Senge, M. O.; Hatscher, S. S.; Wiehe, A.; Dahms, K.; Kelling, A. *J. Am. Chem. Soc.* **2004**, *126*, 13634; (d) Takanami, T.; Wakita, A.; Sawaizumi, A.; Iso, K.; Onodera, H.; Suda, K. *Org. Lett.* **2008**, *10*, 685.
- Inoue, H.; Tsubouchi, H.; Nagaoka, Y.; Tomioka, K. *Tetrahedron* **2002**, *58*, 83.
- (a) Tonogaki, K.; Itami, K.; Yoshida, J. *Org. Lett.* **2006**, *8*, 1419; (b) Tonogaki, K.; Itami, K.; Yoshida, J.-I. *J. Am. Chem. Soc.* **2006**, *128*, 1464.
- Bustelo, E.; Guerot, C.; Hercouet, A.; Carboni, B.; Toupet, L.; Dixneuf, P. H. *J. Am. Chem. Soc.* **2005**, *127*, 11582.
- (a) **Ni8** was synthesised in the optimum yield using the following procedure: To a 20 cm<sup>3</sup> Schlenk tube, **Ni5** (50 mg), PdCl<sub>2</sub>(dppe) (15 mol %, 10 mg) and K<sub>2</sub>CO<sub>3</sub> (107 mg) were added and dried under vacuum before the flask was charged with argon. Dry THF (10 cm<sup>3</sup>) was added and the mixture was degassed via three freeze-pump-thaw cycles. Allenylboronic acid pinacol ester (140 μL, 10-fold excess) was subsequently added, and the flask was sealed and heated to 80 °C overnight. The progress of the reaction was monitored by TLC using CH<sub>2</sub>Cl<sub>2</sub>/*n*-hexane (1:2) as eluent. Upon complete consumption of the starting material, the solvent was removed in vacuo and the residue was chromatographed on silica gel using CH<sub>2</sub>Cl<sub>2</sub>/*n*-hexane (1:2). The first fraction collected was **Ni8** (24 mg, 50%). Analytical data: mp: >300 °C; <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): 9.46 (d, <sup>3</sup>J = 5.0 Hz, 2H, β-H), 8.80 (d, <sup>3</sup>J = 5.0 Hz, 2H, β-H), 8.69 (d, <sup>3</sup>J = 4.90 Hz, 2H, β-H), 8.67 (d, <sup>3</sup>J = 4.90 Hz, 2H, β-H), 8.31 (t, <sup>4</sup>J = 6.8 Hz, 1H, allene-CH), 8.01 (m, 6H, *o*-Ph-H), 7.70 (m, 9H, *m,p*-Ph-H), 5.32 (d, <sup>4</sup>J = 6.8 Hz, 2H, allene-CH<sub>2</sub>); <sup>13</sup>C NMR (150 MHz, CDCl<sub>3</sub>): 216.6, 142.1, 141.9, 141.4, 141.1, 140.5, 140.2, 133.2, 132.2, 132.0, 131.8, 131.7, 130.3, 128.2, 126.5, 125.5, 118.6, 109.3, 92.1; UV/Vis (CH<sub>2</sub>Cl<sub>2</sub>) λ<sub>max</sub> (log ε): 421 (5.41), 535 (4.31), 575 (3.81); HRMS (MALDI) calcd for [M]<sup>+</sup> C<sub>41</sub>H<sub>26</sub>N<sub>4</sub> 632.1511, found 632.1530; IR cm<sup>-1</sup>: ν = 696 (aromatic H), 737 (aromatic H), 1537 (N-H), 1947 (C=C=C), 3054 (aromatic H). (b) **Ni7** was synthesised using **Ni5** (0.074 mmol), PdCl<sub>2</sub>(dppe) (10 mg, 0.018 mmol) and Cs<sub>2</sub>CO<sub>3</sub> (254 mg, 0.78 mmol). The residue was chromatographed on silica gel using CH<sub>2</sub>Cl<sub>2</sub>/*n*-hexane (1:2). The first fraction collected was **Ni7** (29 mg, 61%), mp: >300 °C; <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): 9.63 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.99 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.85 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.84 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.20 (m, 6H, *o*-Ph-H), 7.75 (m, 9H, *m,p*-Ph-H), 5.99 (d, <sup>4</sup>J = 2.6 Hz, 2H, CH<sub>2</sub>), 2.51 (t, <sup>4</sup>J = 2.6 Hz, 1H, acetylene-H); <sup>13</sup>C NMR (150 MHz, CDCl<sub>3</sub>): 144.8, 142.8, 142.4, 142.1, 140.6, 140.5, 133.5, 133.4, 132.5, 132.0, 131.9, 131.5, 127.6, 126.7, 119.7, 119.2, 100.1, 93.4, 79.7, 5.2; UV/Vis (CH<sub>2</sub>Cl<sub>2</sub>) λ<sub>max</sub> (log ε): 420 (5.47), 553 (4.18); HRMS (ESI) calcd. for [M]<sup>+</sup> C<sub>41</sub>H<sub>26</sub>N<sub>4</sub> 632.1449, found 632.1530. (c) **Zn9** was obtained using **Zn5** (0.073 mmol), PdCl<sub>2</sub>(dppe) (10 mg, 0.018 mmol) and K<sub>2</sub>CO<sub>3</sub> (107 mg, 0.78 mmol). The residue was chromatographed on silica gel using CH<sub>2</sub>Cl<sub>2</sub>/*n*-hexane (1:2). The first fraction collected gave **Zn9** (22 mg, 46%), mp: >300 °C; <sup>1</sup>H NMR (400 MHz, CDCl<sub>3</sub>): 9.63 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.99 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.85 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.84 (d, <sup>3</sup>J = 4.7 Hz, 2H, β-H), 8.20 (m, 6H, *o*-Ph-H), 7.75 (m, 9H, *m,p*-Ph-H), 5.99 (d, <sup>4</sup>J = 2.6 Hz, 2H, CH<sub>2</sub>), 2.51 (t, <sup>4</sup>J = 2.6 Hz, 1H, acetylene-H); <sup>13</sup>C NMR (150 MHz, CDCl<sub>3</sub>): 150.1, 150.0, 149.8, 149.6, 143.5, 134.5, 132.5, 131.6, 128.5, 127.1, 126.3, 126.2, 122.9, 120.4, 86.0, 73.1, 29.7;

- UV/Vis (CH<sub>2</sub>Cl<sub>2</sub>) λ<sub>max</sub> (log ε): 420 (5.47), 553 (4.18); LRMS (MALDI) calcd for [M]<sup>+</sup> C<sub>41</sub>H<sub>26</sub>N<sub>4</sub>Zn 638.1449, found 638.1467.
15. Miyaura, N.; Suzuki, A. *Org. Synth.* **1990**, 68, 130.
  16. Newman-Evans, R. H.; Simon, R. J.; Carpenter, B. K. *J. Org. Chem.* **1990**, 55, 695.
  17. Ohno, H. *Chem. Pharm. Bull.* **2005**, 53, 1211.
  18. Marshall, J. A. J. *Org. Chem.* **2007**, 72, 8153.